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(71) Applicant: FINA TECHNOLOGY, INC.
Dallas, Texas 75206 (US)

(72) Inventors:
• Chen, Hong
Alpharetta, Georgia 30005 (US)

• Coffy, Tim J.
Houston, Texas 77062 (US)
• Shamshour, Edward S.
Houston, Texas 77059 (US)

(74) Representative:
Leyder, Francis et al
c/o Fina Research S.A.
Dépt. Brevets
Zone Industrielle C
7181 Seneffe (Feluy) (BE)

(54) **Improved Ziegler-Natta catalysts for olefin polymerization**

(57) A new synthesis of a Ziegler-Natta catalyst uses a multi-step preparation which includes treating a soluble magnesium compound with successively stronger chlorination/titanation reagents. The catalyst may be used in polymerization of olefins, particularly ethylene, to produce a polymer with less fines, larger average fluff particle size and narrow molecular weight distribution. The catalyst has high activity and good hydrogen response.

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Description

BACKGROUND OF THE INVENTION

5 FIELD OF THE INVENTION: This invention relates to catalyst system for the polymerization of olefins, particularly, to a catalyst system comprising a supported Ziegler-Natta catalyst. The catalyst system is used to polymerize olefins, such as ethylene.

DESCRIPTION OF THE PRIOR ART: Polyolefin manufacturing processes typically involve the polymerization of olefin monomer with a Ziegler-Natta type catalyst. Catalyst systems for the polymerization of olefins are well known in the art. Typically, these systems include a Ziegler-Natta type polymerization catalyst component and a co-catalyst, usually an organoaluminum compound. Examples of such catalyst systems are shown in the following U.S. Patents 3,574,138; 4,316,966; and 5,173,465, the disclosures of which are hereby incorporated by reference.

A Ziegler-Natta type polymerization catalyst is basically a complex derived from a halide of a transition metal, for example, titanium or vanadium, with a metal hydride and/or a metal alkyl, typically an organoaluminum compound, as a co-catalyst. The catalyst is usually comprised of a titanium halide supported on a magnesium compound complexed with an alkylaluminum co-catalyst.

SUMMARY OF THE INVENTION

20 Accordingly, an object of this invention is to synthesize a supported Ziegler-Natta catalyst for the polymerization of ethylene.

Also, an object of this invention is to produce a polyolefin having large particle size, a low amount of small particles or "fines" and low wax.

In addition, an object of this invention is to produce a catalyst with high activity and better hydrogen response.

25 These and other objects are accomplished by a catalyst system comprising a supported Ziegler-Natta catalyst used in a polymerization process for polymerizing olefins, especially ethylene.

The present invention provides for a catalyst for polymerization of olefins having high activity and better hydrogen response comprising:

- 30 a) a supported Ziegler-Natta transition metal catalyst component; and
b) an organoaluminum co-catalyst.

The present invention provides for a catalyst component comprising:

- 35 a) a soluble magnesium compound of magnesium dialkoxide of the general formula $Mg(OR)_2$ where R is a hydrocarbyl or substituted hydrocarbyl of 1 to 20 carbon atoms;
b) a mild chlorinating agent;
c) a titanating agent; and
d) an organoaluminum compound.

40 The present invention provides for a process for synthesizing a catalyst component comprising:

- 45 a) synthesizing magnesium di(alkoxide) from magnesium dialkyl and alcohol;
b) adding a mild chlorinating agent;
c) adding a titanating agent;
d) adding a second titanating agent;
e) adding an organoaluminum compound.

50 The present invention also provides a process for the polymerization of olefins using the catalyst system described above to produce a polymer product having a narrow molecular weight distribution, a low amount of small particles and low wax comprising:

- 55 a) selecting a conventional Ziegler-Natta transition metal catalyst component;
b) contacting the catalyst component with an organoaluminum co-catalyst compound;
c) introducing the catalyst system into a polymerization reaction zone containing a monomer under polymerization reaction conditions to form a polymer product; and
e) extracting polymer product from the polymerization reaction zone.

DETAILED DESCRIPTION OF THE INVENTION

The synthesis procedure for Ziegler-Natta type catalysts for the polymerization of olefins is disclosed in U.S. Patent Nos. 3,644,318, the disclosure of which is hereby incorporated. The standard synthesis procedure is:

- a) selecting a magnesium compound;
- b) adding a chlorinating agent;
- c) adding a titanating agent; and
- d) optionally, adding a preactivating agent.

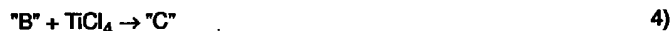
The chlorinating agent and the titanating agent may be the same compound.



The present invention modifies the synthesis procedure, and thus modifies the catalyst and the polymer product. The present invention is characterized by the following:

- 1) a soluble magnesium compound;
- 2) a mild chlorinating agent;
- 3) simultaneous chlorination and titanation steps with a mild reagent;
- 4) a second chlorination/titanation with a stronger reagent;
- 5) a preactivation step.

A proposed mechanism for the modified synthesis procedure is as follows:



While the exact composition of "A" is unknown, it is believed that it contains a partially chlorinated magnesium compound, one example of which may be $\text{ClMg(OR}'')$. The first chlorination/titanation produces a catalyst ("B") which is probably a complex of chlorinated and partially chlorinated magnesium and titanium compounds and may possibly be represented by $(\text{MgCl}_2)_y \cdot (\text{TiCl}_x(\text{OR}')_{4-x})_z$. The second chlorination/titanation produces a catalyst ("C") which is also probably a complex of chlorinated and partially chlorinated magnesium and titanium compounds but different from "B" and may possibly be represented by $(\text{MgCl}_2)_y \cdot (\text{TiCl}_x(\text{OR}')_{4-x})_z$. It is expected that the level of chlorination of "C" would be greater than that of "B". This greater level of chlorination would produce a different complex of different compounds. While this description of the reaction products offers the most probable explanation of the chemistry at this time, the invention as described in the claims is not limited by this theoretical mechanism.

The soluble magnesium compound is preferably a non-reducing compound such as magnesium dialkoxide of the general formula $\text{Mg(OR}'')_2$ where R'' is a hydrocarbyl or substituted hydrocarbyl of 1 to 20 carbon atoms. A non-reducing compound has the advantage of forming MgCl_2 instead of insoluble Ti^{+3} species formed by reduction of compounds such as MgRR' which tend to form catalysts having a broad particle size distribution. In addition, $\text{Mg(OR}'')_2$ is less reactive than MgRR' and the chlorination with a mild chlorinating agent, followed by a simultaneous chlorination/titanation with a mild reagent and a second chlorination/titanation with a stronger reagent are gradual and successively stronger reactions which may result in more uniform product, i.e., larger catalyst particles and better catalyst particle size control.

Magnesium dialkoxide, such as magnesium di(2-ethylhexoxide), may be produced by reacting an alkyl magnesium compound (MgRR'), such as butyl ethyl magnesium (BEM), with an alcohol (ROH), such as 2-ethylhexanol.



In the case of BEM, RH and R'H are butane and ethane, respectively. The reaction takes place at room temperature and the reactants form a solution.

The magnesium dialkyl $[MgRR']$ may be any magnesium dialkyl where R and R' are alkyl groups of 1-10 carbon atoms. R and R' may be the same or different. Examples of the magnesium dialkyl are magnesium diethyl, magnesium dipropyl, magnesium dibutyl, butylethylmagnesium, etc. Butylethylmagnesium (BEM) is the preferred magnesium dialkyl.

The alcohol may be any alcohol of the general formula $R''OH$ where R'' is an alkyl group of 4-20 carbon atoms. The alcohol may be linear or branched. Examples of the alcohol are butanol, isobutanol, 2-ethylhexanol, etc. The preferred alcohol is 2-ethylhexanol.

Alkyl magnesium compounds are highly associative due to electron-deficient bonding which results in a high molecular weight species which is very viscous in solution. This high viscosity may be reduced by the addition of an aluminum alkyl, such as triethylaluminum, which disrupts the association between the individual alkyl magnesium molecules. The preferred ratio of alkyl aluminum to magnesium is 0.001:1 to 1:1, more preferably 0.01 to 0.1:1 and most preferably 0.03:1 to 0.05:1. In addition, an electron donor such as an ether, e.g., diisooamyl ether (DIAE) may be used to further reduce the viscosity of the alkyl magnesium. The preferred ratio of electron donor to magnesium is 0:1 to 10:1 and more preferably is 0.1:1 to 1:1.

The chlorinating agent is preferably a monochloride compound which only partially chlorinates the magnesium alkoxide. The chlorinating agent is of the general formula $ClAR''_x$ or $ClAOR''_x$, where A is a nonreducing oxyphilic compound which is capable of exchanging one chloride for an alkoxide, R'' is alkyl and x is the valence of A minus 1. Examples of A are titanium, silicon, aluminum, carbon, tin and germanium, most preferred of which is titanium and silicon wherein x is 3. Examples of R'' are methyl, ethyl, propyl, isopropyl and the like having 2-6 carbon atoms. Examples of a chlorinating agent effective in the present invention are $ClTi(O^iPr)_3$ and $ClSi(Me)_3$.

The chlorinating/titanating agent is preferably a tetra-substituted titanium compound with all four substituents being the same and the substituents being a halide or an alkoxide or phenoxide with 2 to 10 carbon atoms, such as $TiCl_4$ or $Ti(OR'')_4$. The chlorinating/titanating agent may be a single compound or a combination of compounds. The synthesis provides an active catalyst after the first chlorination/titanation; however, chlorination/titanation preferably is done twice, is a different compound or combination of compounds at each addition and is a stronger chlorination/titanation at the second addition.

The first chlorination/titanation agent is preferably a mild titanating agent which is preferable a blend of a titanium halide and a titanium alkoxide. More preferably, the first chlorinating/titanating agent is a blend of $TiCl_4$ and $Ti(OBu)_4$ in a range from 0.5:1 to 6:1 $TiCl_4/Ti(OBu)_4$, most preferably from 2:1 to 3:1. The ratio of titanium to magnesium in the first titanation is preferably 3:1. It is believed that the blend of titanium halide and titanium alkoxide react to form a titanium alkoxyhalide, $Ti(OR)_aX_b$, where OR and X are an alkoxide and halide, respectively, a + b is the valence of titanium which is typically 4 and both a and b may be fractional, e.g., a=2.5 and b=1.5.

In the alternative, the first chlorination/titanation agent may be a single compound. Examples of the first chlorination/titanation agent as a single compound are $Ti(OC_2H_5)_3Cl$, $Ti(OC_2H_5)_2Cl_2$, $Ti(OC_3H_7)_2Cl_2$, $Ti(OC_3H_7)_3Cl$, $Ti(OC_4H_9)_3Cl$, $Ti(OC_6H_{13})_2Cl_2$, $Ti(OC_2H_5)_2Br_2$ and $Ti(OC_{12}H_{25})Cl_3$.

Optionally, an electron donor may be added with the first mild chlorination/titanation agent. It is believed that the addition of an electron donor at this step may result in better particle size control. The electron donor is preferably an alkylsilylalkoxide of the general formula $RSi(OR')_3$, e.g., methylsilyltriethoxide $[MeSi(OEt)_3]$, where R and R' are alkyls with 1-5 carbon atoms and may be the same or different.

The second stronger chlorinating/titanating agent is preferably a titanium halide, more preferably titanium tetrachloride $[TiCl_4]$. The range of titanium to magnesium in the second chlorination/titanation is 0:1 to 2:1 and is preferably 0.5:1 to 1:1.

The preactivating agent is preferably an organoaluminum compound. The organoaluminum preactivating agent is preferably an aluminum alkyl of the formula AlR^A_3 where R^A is an alkyl having 1-8 carbon atoms or a halide, R' being the same or different and at least one R is an alkyl. The organoaluminum preactivating agent is preferably a trialkyl aluminum, such as trimethyl aluminum (TMA), triethyl aluminum (TEAL) and triisobutyl aluminum (TiBAI). The preferred preactivating agent is TEAL. The ratio of Al to titanium is in the range from 0.1:1 to 2:1 and preferably is 0.5:1 to 1.2:1 and more preferably is approximately 0.8:1.

The specific synthesis procedure used was:

- synthesizing soluble magnesium dialkoxide [magnesium di(2-ethyl-hexoxide)] from magnesium dialkyl [butylethylmagnesium] and an alcohol [2-ethylhexanol], optionally adding an aluminum alkyl (triethylaluminum) for viscosity control;
- adding a mild monochloro-chlorinating agent $[ClTi(O^iPr)_3]$;
- adding a first mild chlorinating/titanating agent blend of titanium tetrahalide and titanium tetra-alkoxide $[TiCl_4/Ti(OBu)_4]$;
- adding a second stronger chlorinating/titanating agent of a titanium tetrahalide $[TiCl_4]$;
- adding an alkylaluminum [TEAL] to preactivate the catalyst.

The conventional supported Ziegler-Natta transition metal compound catalyst component as described above may be used in the polymerization of olefins, particularly ethylene. The transition metal compound is preferably of the general formula MR^1_x where M is the metal, R^1 is a halogen or a hydrocarbyloxy and x is the valence of the metal. Preferably, M is a Group IVB metal, more preferably titanium. Preferably, R^1 is chlorine, bromine, an alkoxy or a phenoxy, more preferably chlorine. Illustrative examples of the transition metal compounds are $TiCl_4$, $TiBr_4$, $Ti(OC_2H_5)_3Cl$, $Ti(OC_2H_5)_3Cl$, $Ti(OC_3H_7)_2Cl_2$, $Ti(OC_4H_9)_3Cl$, $Ti(OC_6H_{13})_2Cl_2$, $Ti(OC_8H_{17})_3Cl$, $Ti(OC_2H_5)_2Br_2$ and $Ti(OC_{12}H_{25})Cl_3$. Mixtures of the transition metal compounds may be used. No restriction on the number of transition metal compounds is made as long as at least one transition metal compound is present.

The support should be an inert solid which is chemically unreactive with any of the components of the conventional Ziegler-Natta catalyst. The support is preferably a magnesium compound. Examples of the magnesium compounds which are to be used to provide a support for the catalyst component are magnesium halides, dialkoxymagnesiums, alkoxymagnesium halides and carboxylates of magnesium. The preferred magnesium compound is a magnesium chloride ($MgCl_2$).

The catalyst is activated with an organoaluminum cocatalyst. The organoaluminum co-catalyst is preferably an aluminum alkyl of the formula AlR^A_3 where R^A is an alkyl having 1-8 carbon atoms or a halide, R^A being the same or different and at least one R^A is an alkyl. The organoaluminum co-catalyst may be an aluminum trialkyl, an aluminum dialkyl halide or an aluminum alkyl dihalide. More preferably, the organoaluminum co-catalyst is trialkyl aluminum, such as trimethyl aluminum (TMA), triethyl aluminum (TEAL) and triisobutyl aluminum (TIBAL). The preferred aluminum alkyl is TEAL.

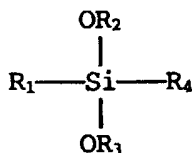
A Ziegler-Natta catalyst may be pre-polymerized to improve the performance of the catalyst. Generally, a prepolymerization process is effected by contacting a small amount of monomer with the catalyst after the catalyst has been contacted with the co-catalyst. A pre-polymerization process is described in U.S. Patent Nos. 5,106,804; 5,153,158; and 5,594,071, hereby incorporated by reference.

The catalyst may be used in any known process for the homopolymerization or copolymerization of olefins. The polymerization process may be bulk, slurry or gas phase. It is preferred to use a catalyst synthesized above in a slurry phase polymerization in temperature range of 50-100°C, preferably 50-120°C, more preferably 70-80°C, and a pressure range of 50-800 psi, preferably 80-600 psi, more preferably 100-150 psi. The olefin monomer may be introduced into the polymerization reaction zone in a diluent which is a nonreactive heat transfer agent which is liquid at the reaction conditions. Examples of such a diluent are hexane and isobutane. For the copolymerization of ethylene with another alpha-olefin, such as butene, the second alpha-olefin may be present at 0.01-20 mole percent, preferably 0.02-1 mole percent and more preferably about 0.0625 mole percent.

An internal electron donor may be used in the formation reaction of the catalyst during the chlorination or chlorination/titanation steps. Compounds suitable as internal electron-donors for preparing conventional supported Ziegler-Natta catalyst components include ethers, ketones, lactones, electron donors compounds with N, P and/or S atoms and specific classes of esters. Particularly suitable are the esters of phthalic acid, such as diisobutyl, dioctyl, diphenyl and benzylbutylphthalate; esters of malonic acid, such as diisobutyl and diethylmalonate; alkyl and arylpivalates; alkyl, cycloalkyl and arylmaleates; alkyl and aryl carbonates such as diisobutyl, ethyl-phenyl and diphenylcarbonate; succinic acid esters, such as mono and diethyl succinate.

External donors which may be utilized in the preparation of a catalyst according to the present invention include organosilane compounds such as alkoxysilanes of general formula $SiR_m(OR')_{4-m}$ where R is selected from the group consisting of an alkyl group, a cycloalkyl group, an aryl group and a vinyl group; R' is an alkyl group; and m is 0-3, wherein R may be identical with R' ; when m is 0, 1 or 2, the R' groups may be identical or different; and when m is 2 or 3, the R groups may be identical or different.

Preferably, the external donor of the present invention is selected from a silane compound of the following formula:



wherein R_1 and R_4 are both an alkyl or cycloalkyl group containing a primary, secondary or tertiary carbon atom attached to the silicon, R_1 and R_4 being the same or different; R_2 and R_3 are alkyl or aryl groups. R_1 may be methyl, isopropyl, cyclopentyl, cyclohexyl or t-butyl; R_2 and R_3 may be methyl, ethyl, propyl, or butyl groups and not necessarily the same; and R_4 may also methyl, isopropyl, cyclopentyl, cyclohexyl or t-butyl. Specific external donors are cyclohexylmethyldimethoxy silane (CMDS), diisopropyldimethoxysilane (DIDS) cyclohexylisopropyl dimethoxysilane (CIDS),

dicyclopentylidimethoxysilane (CPDS) or di-*t*-butyl dimethoxysilane (DTDS).

The invention having been generally described, the following examples are given as particular embodiments of the invention and to demonstrate the practice and advantages thereof. It is understood that the examples are given by way of illustration and are not intended to limit the specification or the claims to follow in any manner.

The following parameters were varied in the Examples below:

1. Different chlorination/precipitation reagent
2. Catalyst preparation with and without using the monochloride
3. Use of an alternative monochloride
4. Use of diisooamyl ether (DIAE) or Me(Si(OEt)₃) as an electron donor

Catalyst preparation

EXAMPLE 1

50 mmole of BEM was weighed into a 100 ml measuring cylinder in the nitrogen box. 1.5 mmole of TEAL was added to the BEM solution. The BEM/TEAL mixture was diluted to 100 ml with heptane, and transferred to a 1L 5-neck flask. The viscosity of this mixture was much lower than BEM itself. The measuring cylinder was rinsed twice with 50 ml portions of heptane, which was also transferred into the 1L flask. The solution was stirred at room temperature at 200 rpm.

104.5 mmole 2-Et-HexOH was weighed into a 50 ml measuring cylinder and diluted to 50 ml with heptane. The alcohol solution was transferred to the 60 ml addition funnel on the 1L flask, and added dropwise to the BEM/TEAL solution at RT over 50 mins. The addition funnel was rinsed with 50 ml heptane. The reaction mixture was stirred at room temperature for another hour.

50 mmole ClTi(O^{*i*}Pr)₃ (hexane solution) was weighed into a 100 ml measuring cylinder and transferred to the 60 ml addition funnel. The measuring cylinder was rinsed with 10 ml heptane. The ClTi(O^{*i*}Pr)₃ solution was added dropwise to the flask at room temperature over 70 mins. The reaction mixture was stirred at room temperature for another hour.

19.0 g TiCl₄ was weighed into a 100 ml measuring cylinder and diluted with heptane to 100 ml. The TiCl₄ solution was transferred to the addition funnel and added dropwise to the flask at room temperature over 55 mins. The reaction mixture was stirred at room temperature for one hour. The reaction mixture was then allowed to settle. The supernatant was decanted. The solid was washed four times with 200 ml portions of heptane at 40°C. The solid was dried *in vacuo* at 40°C for 2 hrs. (white powder, ca. 9.6 g).

EXAMPLE 2

50 mmole of BEM was weighed into a 100 ml measuring cylinder and 1.5 mmole of TEAL was added to the BEM solution. The BEM/TEAL mixture was diluted to 100 ml with heptane. The BEM solution was transferred to a 1 L flask equipped with a thermometer, a mechanical stirrer, a 60 ml addition funnel and a condenser with a gas inlet. The measuring cylinder was rinsed with 5 ml heptane. The solution was agitated at 200 rpm at room temperature.

104.5 mmole of 2-Et-HexOH was weighed into a 50 ml measuring cylinder and diluted to 50 ml with heptane. The alcohol solution was transferred to the 60 ml addition funnel on the 1 L flask. The cylinder was rinsed with 10 ml heptane. The alcohol solution was added dropwise to the BEM/TEAL solution at room temperature over 33 mins. The addition funnel was rinsed with 50 ml heptane. The reaction mixture was stirred at room temperature for another hour.

50 mmoles of ClTi(O^{*i*}Pr)₃ (hexane solution) was weighed into a 100 ml measuring cylinder and transferred to the 60 ml addition funnel on the 1 L flask. The measuring cylinder was rinsed with 10 ml heptane. The ClTi(O^{*i*}Pr)₃ solution was added dropwise to the flask at room temperature over 46 mins. The addition funnel was rinsed with 10 ml heptane. The reaction mixture was stirred at room temperature for two hours.

19.0 g TiCl₄ and 17.0 g Ti(OBu)₄ were weighed and mixed in a 100 ml measuring cylinder in the nitrogen box. The mixture was diluted with heptane to 100 ml, and transferred to the addition funnel and added dropwise to the flask at room temperature over 1.5 hrs. The reaction mixture was stirred at room temperature for three hours. The reaction mixture was then allowed to settle. The supernatant was decanted. The white solid was washed 4 times with 200 ml portions of heptane.

9.5 g TiCl₄ was weighed into a 100 ml measuring cylinder and diluted to 50 ml with heptane. The TiCl₄ solution was transferred to the 60 ml addition funnel and added dropwise to the flask at room temperature over 35 mins. The reaction mixture was stirred at room temperature for another two hours. The reaction mixture was allowed to settle. The solid was washed four times with 150 ml portions of heptane. The solid was dried *in vacuo* at room temperature for 1.5 hrs. (off-white powder, ca. 5.5 g).

EXAMPLE 3

The procedure of Example 2 was used to prepare a solution (A) from BEM/TEAL + 2-Et-HexOH + $\text{CITi}(\text{O}^i\text{Pr})_3$.

17.0 g $\text{Ti}(\text{OBu})_4$ was weighed into a 50 ml measuring cylinder in the nitrogen box. 1.78 g $\text{MeSi}(\text{OEt})_3$ was added.

5 The mixture was diluted with heptane to 50 ml. This solution was transferred to a 1 L 5-neck flask with standard equipment. The solution was stirred at room temperature at 150 rpm. Then 19.0 g TiCl_4 diluted to 50 ml in heptane was added to the $\text{Ti}(\text{OBu})_4/\text{MeSi}(\text{OEt})_3$ solution. This solution (B) was stirred at room temperature for 35 mins.

The solution A prepared above was transferred to the 60 ml addition funnel on the 1 L flask, and added dropwise to the flask at room temperature over 1 hr 50 mins. The reaction mixture was stirred at ambient temperature for another 10 30 mins and then allowed to settle at room temperature.

The supernatant was decanted. The solid was washed three times with 200 ml portions of hexane. 200 ml fresh hexane was added to the flask and the slurry was agitated at 150 rpm at room temperature.

9.5 g TiCl_4 was weighed into a 50 ml measuring cylinder and diluted to 50 ml with hexane. The TiCl_4 solution was transferred to the 60 ml addition funnel on the 1 L flask, and added dropwise to the flask in 25 mins. The reaction mixture was stirred at room temperature for another two hours. The reaction mixture was then allowed to settle. The supernatant was decanted. The solid was washed 4 times with 200 ml portions of hexane. The solid was dried *in vacuo* at room temperature for 2 hrs. The yield of the solid was ca. 5.1 g.

EXAMPLE 4

20

The procedure to prepare this catalyst is the same as that for Example 3, except that the mixture of $\text{TiCl}_4/\text{Ti}(\text{OBu})_4/\text{MeSi}(\text{OEt})_3$ (2/1/0.2) was added to the solution "A" prepared from BEM/TEAL (1/0.03) + 2-Et-HexOH (2) + $\text{CITi}(\text{O}^i\text{Pr})_3$ (1). The resultant white solid was dried *in vacuo* at room temperature for 1.5 hours (ca. 4.2 g).

EXAMPLE 5

25 The $\text{Mg}(\text{OR})_2$ solution was prepared with BEM/TEAL (50 mmole/1.5 mmole) and 2-Et-HexOH (104.5 mmole) as in Example 2. The viscous solution was stirred at ambient temperature at 150 rpm.

104.5 mmole of $\text{CITi}(\text{O}^i\text{Pr})_3$ (hexane solution) was weighed into a 100 ml measuring cylinder and then transferred 30 to the addition funnel on the 1 L flask. The $\text{CITi}(\text{O}^i\text{Pr})_3$ solution was added dropwise to the flask over 1.5 hrs. The reaction mixture was stirred at room temperature for another 1.5 hrs. The reaction mixture was allowed to settle. The settling of the solid was very slow. The supernatant was decanted and the solid was re-slurried in 150 ml heptane.

A solution of $\text{TiCl}_4/\text{Ti}(\text{OBu})_4/\text{MeSi}(\text{OEt})_3$ (100 mmole/50 mmole/10 mmole) was prepared in the same way as for the preparation of Example 3. This solution was transferred to the 60 ml addition funnel on the 1 L flask, and added 35 dropwise to the flask at room temperature over a period of 2 hrs. The reaction mixture was stirred at room temperature for another hour. The reaction mixture was allowed to settle. The settling of the solid was much faster than before the TiCl_4 addition. The supernatant was decanted. The solid was washed four times with 150 ml portions of heptane.

9.5 g TiCl_4 was weighed into a 50 ml measuring cylinder and diluted to 40 ml with heptane. The TiCl_4 solution was transferred to the 60 ml addition funnel. The cylinder was rinsed with 10 ml heptane. The TiCl_4 solution was added drop- 40 wise to the flask at room temperature over 35 mins. The reaction mixture was stirred at room temperature for another hour. The reaction mixture was allowed to settle. The supernatant was decanted. The solid was washed four times with 150 ml portions of heptane. The solid was dried *in vacuo* at room temperature for 2 hrs. The yield of the dried catalyst (white, static) was 1.3 g.

EXAMPLE 6

45 The solution of $\text{Mg}(\text{OR})_2$ was prepared with BEM/TEAL (50 mmole/1.5 mmole) and 2-Et-HexOH (104.5 mmole) using the procedure of Example 2.

19.0 g TiCl_4 was weighed into a 50 ml measuring cylinder and diluted to 20 ml with heptane. The TiCl_4 solution was 50 cannulated into a 100 ml flask containing 17.0 g $\text{Ti}(\text{OBu})_4$. The orange-brown mixture was stirred at room temperature for 20 mins., then transferred to the 60 ml addition funnel on the 1 L flask. The 100 ml flask was rinsed with a small amount of heptane. The $\text{TiCl}_4/\text{Ti}(\text{OBu})_4$ solution was added dropwise to the 1 L flask at room temperature over 55 mins. The reaction mixture was stirred at room temperature for another hour. The reaction mixture was then allowed to settle. The settling was very slow (>1.5 hrs). The clear, light yellow supernatant was decanted. The white solid was washed 55 three times with 150 ml portions of heptane. The solid was dried *in vacuo* at room temperature for four hours. The solid was yellow, lumpy. The flask was then heated to 70 °C and the solid was dried for another hour. The resultant solid was more powdery but still lumpy. The yield of the solid was ca. 8.3 g.

EXAMPLE 7

The preparation of $\text{Mg}(\text{OR})_2$ solution was the same as Example 6 except that 1.58 g DIAE was added to the mixture of BEM and TEAL before the reaction. The viscosity of the BEM/TEAL solution was drastically reduced upon the addition of DIAE. The reaction was carried out at room temperature with 150 rpm agitation.

A mixture of 17.0 g $\text{Ti}(\text{OBu})_4$ and 19.0 g TiCl_4 was prepared in the same manner as in Example 2. This solution was added dropwise to the reaction flask at room temperature over 55 mins. After the addition was completed, the reaction mixture was stirred at room temperature for another hour. The reaction mixture was then allowed to settle. The flask was slowly heated to 50 °C during the settling. The supernatant was decanted. The white solid was washed four times with 150 ml portions of heptane at 45-50 °C. The solid was dried *in vacuo* at 50-55 °C for 3 hrs. The solid (8.7 g) was more powdery than that of Example 6 but was still lumpy.

EXAMPLE 8

The preparation of $\text{Mg}(\text{OR})_2$ solution was the same as for Example 7. The reaction was carried out at room temperature with 150 rpm agitation.

50 mmole of $\text{CITi}(\text{O}^i\text{Pr})_3$ was weighed into a 100 ml measuring cylinder and transferred to the 60 ml addition funnel. The measuring cylinder was rinsed with 10 ml heptane. The $\text{CITi}(\text{O}^i\text{Pr})_3$ solution was added dropwise to the flask at room temperature over 40 mins. The addition funnel was rinsed with 5 ml heptane. The reaction mixture was stirred at room temperature for another hour.

A mixture of 17.0 g $\text{Ti}(\text{OBu})_4$ and 19.0 g TiCl_4 was prepared as in Example 6. The $\text{TiCl}_4/\text{Ti}(\text{OBu})_4$ solution was transferred to the addition funnel and added dropwise to the flask at room temperature over 75 minutes. The reaction mixture was stirred at room temperature for one hour. The reaction mixture was then allowed to settle. The supernatant was decanted. The solid was washed twice with 150 ml portions of heptane, then re-slurried in 150 ml heptane.

9.5 g TiCl_4 was weighed into a 50 ml measuring cylinder and diluted to 60 ml with heptane. The TiCl_4 solution was added dropwise to the flask at room temperature over 40 mins. The reaction mixture was stirred at room temperature for another hour. The reaction mixture was then allowed to settle. The supernatant was decanted. The solid was washed four times with 150 ml portions of heptane. The solid was dried *in vacuo* at room temperature for 2 hrs. The yield of the solid white powder was ca. 9.6 g.

EXAMPLE 9

The solution of $\text{Mg}(\text{OR})_2$ was prepared as in Example 2 using BEM/TEAL and 2-Et-HexOH. The reaction was carried out at room temperature under agitation of 150 rpm.

13.5 g ClSiMe_3 was weighed into a 50 ml measuring cylinder and diluted with heptane to 40 ml. The solution was transferred to the 60 ml addition funnel on the 1 L flask. The cylinder was rinsed with 10 ml heptane. The ClSiMe_3 solution was added dropwise to the 1 L flask at room temperature over 30 mins. The reaction mixture was stirred at room temperature for another 1 h and 45 mins.

A mixture of 17.0 g $\text{Ti}(\text{OBu})_4$ and 19.0 g TiCl_4 was prepared as for Example 6. This mixture was added dropwise to the 1 L flask at room temperature over 70 mins. The reaction mixture was stirred at room temperature for another hour. The reaction mixture was allowed to settle. The settling was very slow. The supernatant was decanted. The solid was washed with 150 ml heptane at room temperature. The solid was dried *in vacuo* for 2 hrs. The catalyst was very lumpy with a yellow color. It was heated to 70 °C and dried for another hour. The solid became more powdery but still lumpy. The yield of the solid was 9.7 g.

EXAMPLE 10

The reaction was carried out as in Example 9, except that the ClSiMe_3 solution was added to the $\text{Mg}(\text{OR})_2$ solution at 50 °C instead of room temperature.

The addition of the $\text{TiCl}_4/\text{Ti}(\text{OBu})_4$ solution was carried out room temperature over 75 minutes. The reaction mixture was stirred at room temperature for one hour after the addition of the $\text{TiCl}_4/\text{Ti}(\text{OBu})_4$ solution was completed. Then the reaction mixture was allowed to settle. The supernatant was decanted and the solid was washed twice with 150 ml portions of heptane. The solid was then re-slurried in 150 ml fresh heptane.

A second titration was performed at room temperature using a 50 ml heptane solution containing 9.5 g TiCl_4 . The reaction mixture was stirred at room temperature for one hour after the TiCl_4 addition was completed. Then the reaction mixture was allowed to settle. The supernatant was decanted. The solid was washed four times with 150 ml portions of heptane. The solid was dried *in vacuo* for 1 h. The catalyst was heated to 55-60 °C and dried for another 2.5 hrs. The final solid was off-white and a little lumpy.

EXAMPLE 11

10 g $\text{Mg}(\text{OEt})_2$ was added to a 3-neck 500 mL flask fitted with a 60 mL dropping funnel, a condenser with gas inlet and a septum. 200 mL heptane was added and the slurry was stirred while heating to the reaction temperature (85°C).
 24 mL TiCl_4 was added to the stirred slurry over 15 minutes. The reaction mixture was stirred at 85°C for 5.5 hrs. The reaction mixture was cooled to 75°C and allowed to settle. The supernatant was decanted. The solid was washed four times each with ~100 mL heptane at 70°C. The solid was slurried in ~100 mL dry octane and heated to 120°C for 18 hours. The slurry was cooled to 70°C and then the solid was allowed to settle. The supernatant was decanted. The solid was dried at 70°C for 2 hours in vacuo. The yield was ca. 14g.

Polymerization

A four liter reactor was fitted with four mixing baffles with two opposed pitch mixing propellers. Ethylene and hydrogen were introduced to the reactor vessel via mass flow controllers while a back-pressure regulator keeps the internal reaction pressure constant. The reaction temperature was maintained (in the reactor jacket) by steam and cold water using a valve linked to a controller.

Hexane was used as diluent and TEAL is used as cocatalyst for all polymerizations.

Polymerizations were carried out under the following conditions:

Polymerization Conditions

Temperature 80°C
 Reaction Time 60 minutes
 Catalyst 5-10 mg

TABLE I

Example	Yield(g)	Productivity (g/g · hr)	% fines (<125μm)	Average Catalyst Particle Size (microns)
1.	283	21,000	34.5	10.6
2.	280	21,000	20.5	15.5
3.	369	28,000	17	19.6
4.	327	25,000	5.4	15.8
5.	149	25,000	10	12.2
6.	410	31,000	19.3	10.3
7.	301	23,000	17.9	12.5
8.	375	28,000	1.0	11.0
9.	49	4,000	7.1	21.8
10.	330	20,000	32.7	26.5
11.	260	26,000	5.0	9.0

An improved catalyst synthesis is as follows: BEM, TEAL and DIAE are added in a molar ratio of 1:0.03:0.6 in a solution at room temperature. An amount of 2-Et-HexOH equal to 2.09 equivalents is added to the BEM/TEAL/DIAE solution at room temperature to form magnesium dialkoxide. One equivalent of $\text{ClTi}(\text{O}^i\text{Pr})_3$ is added to the solution at room temperature. A mixture of TiCl_4 and $\text{Ti}(\text{OBu})_4$ as a titanating agent in a molar ratio of 2:1 is added to the solution at room temperature to form an initial catalyst. A second titration with TiCl_4 in the amount of one equivalent is performed on the catalyst solution.

The following parameters were varied in the Examples below:

1. The concentration of the $\text{TiCl}_4/\text{Ti}(\text{OBu})_4$ mixture - Example 12
2. The amount of TiCl_4 used for the second titration - Example 13
3. The titration temperature - Examples 14 and 17

4. The amount of DIAE - Examples 15 and 16
5. Heat treatment after titration - Examples 18 and 20
6. The preactivation of the catalyst - Example 19
7. The impact of TEAL during synthesis - Example 21
8. The slurry concentration - Example 22

EXAMPLE 12

The first two steps of the synthesis, i.e., the preparation of $\text{Mg}(\text{OR})_2$ and its reaction with $\text{ClTi}(\text{O}^i\text{Pr})_3$ were carried out using the same procedure as for Example 8.

19.0 g TiCl_4 was weighed into a 50 ml measuring cylinder and diluted to 50 ml with heptane. The TiCl_4 solution was cannulated into a 200 ml flask containing 17.0 g $\text{Ti}(\text{OBu})_4$. The cylinder was rinsed with 20 ml heptane. The orange-brown mixture was stirred at room temperature for 35 mins. The solution (ca. 86 ml) was transferred to the 60 ml addition funnel on the 1 L flask (60 ml first), and added dropwise to the flask at room temperature over 92 minutes. The reaction mixture was stirred at room temperature for one hour, then allowed to settle. The supernatant was decanted. The solid was washed twice with 180 ml portions of heptane, then re-slurried in 180 ml heptane.

9.5 g TiCl_4 was weighed into a 50 ml measuring cylinder and diluted to 60 ml with heptane. The TiCl_4 solution was added dropwise to the flask at room temperature over 35 mins. The reaction mixture was stirred at room temperature for another hour and then was allowed to settle. The supernatant was decanted. The solid was washed four times with 200 ml portions of heptane. The solid was dried *in vacuo* at room temperature for 2 hrs. The yield of the white solid was ca. 4.8 g.

EXAMPLE 13

The procedure to prepare this catalyst was the same as that for catalyst Example 12, except that 4.75 g TiCl_4 (0.5 eq.) instead of 9.5 g TiCl_4 (1 eq.) was used for the second titration. The yield of the solid catalyst was ca. 5.2 g.

EXAMPLE 14

The procedures to prepare this catalyst is the same as that for Example 13, except that the first titration using $\text{TiCl}_4/\text{Ti}(\text{OBu})_4$ was carried out at 0 °C instead of the ambient temperature. The rate of precipitation seemed to be slower than the room temperature reaction, but the particle size was very small. The ice bath was removed once the addition of $\text{TiCl}_4/\text{Ti}(\text{OBu})_4$ was completed, and the reaction mixture was stirred at ambient temperature for one hour. The catalyst settling was much slower than the catalyst prepared at room temperature. The second titration was carried out in the same way as for Example 13. The resultant white solid was dried *in vacuo* at room temperature for 2.5 hours (ca. 6.7 g). The dried solid catalyst was slightly static and looked very fluffy.

EXAMPLE 15

The catalyst of this example was prepared using the same procedures as for Example 13 except that 3.16 g DIAE (0.4 eq. to Mg) was used as opposed to 1.58 g (0.2 eq.) DIAE. The reaction proceeded similarly to Example 13. The yield of the solid was 5.0 g.

EXAMPLE 16

The catalyst of this example was prepared using the same procedures as for Example 13, except that 4.74 g DIAE (0.6 eq. to Mg) was used as opposed to 1.58 g (0.2 eq.) DIAE. The reaction proceeded similarly to Example 13. The yield of the solid was 5.1 g.

EXAMPLE 17

The preparation of the catalyst of this example followed the same procedures as for Example 16, except that the first titration was carried out at 50 °C instead of the ambient temperature. No preactivation was performed on this catalyst. The yield of the solid catalyst was 7.7 g.

EXAMPLE 18

The first few steps of the catalyst synthesis was the same as that for Example 16. However, after the first titration

was completed and the reaction mixture was stirred at room temperature for one hour, the reaction mixture was heated to 50°C and stirred for one hour. The rest of the synthesis proceeded as per Example 16. The yield of the dried solid catalyst was 4.7 g.

5 **EXAMPLE 19**

The first few steps of the reaction were carried out in the same way as Example 16.

After the second titration 7.83 mmol TEAL was weighed into a 100 ml measuring cylinder and diluted to 60 ml with heptane. The TEAL solution was added dropwise to the 1 L flask at room temperature over 34 mins. The reaction mixture was stirred at ambient temperature for another hour. The reaction mixture was allowed to settle. The supernatant was decanted and the brown solid was washed four times with 200 ml portions of heptane. The solid was dried in vacuo at room temperature for 2.5 hrs (7.4 g).

EXAMPLE 20

15 The catalyst synthesis followed the same procedures as that for Example 17, except that a preactivation was performed after the second titration similar to Example 19 using 7.24 mmole TEAL at room temperature. The yield of the dried, brown solid catalyst was 6.6 g.

20 **EXAMPLE 21**

50 mmol of BEM in heptane solution was weighed into a 100 ml measuring cylinder. 4.74 g DIAE (30 mmole) was added to the BEM solution. The mixture was diluted to 80 ml with heptane, and transferred to a 1 L 5-neck flask with standard set-up. The solution was stirred at room temperature at 150 rpm.

25 13.00 g 2-Et-HexOH was weighed into a 50 ml measuring cylinder and diluted to 50 ml with heptane. The alcohol solution was transferred to the 60 ml addition funnel on the 1 L flask. The cylinder was rinsed with 10 ml heptane. The 2-Et-HexOH solution was added dropwise to the 1 L flask at room temperature over 30 minutes. At the end of the addition, the reaction mixture became a highly viscous gel.

30 1.5 mmole TEAL was weighed into a 20 ml Wheaton bottle and diluted to 10 ml with heptane. The TEAL solution was cannulated into the 1 L flask. The viscosity of the reaction mixture remained high and still a lot of gel was sticking on the wall. The reaction mixture was stirred at room temperature for 20 mins, and no significant improvement in viscosity was observed.

0.59 g 2-Et-HexOH was weighed into a 20 ml Wheaton bottle and diluted to 10 ml with heptane. The solution was transferred to the 60 ml addition funnel. The Wheaton bottle was rinsed with 10 ml heptane. The viscosity of the reaction mixture decreased dramatically and became normal.

35 The rest of the catalyst preparation, i.e., addition of $\text{CITi}(\text{O}^i\text{Pr})_3$, the first titration with $\text{TiCl}_4/\text{Ti}(\text{O}^i\text{Bu})_4$ and the second titration with TiCl_4 was carried out in as usual. The yield of the solid catalyst was 6.2 g.

EXAMPLE 22

40 100 mmol of BEM in heptane solution was weighed into a 250 ml measuring cylinder. 1.38 g TEAL (24.8% in heptane) and 9.48 g DIAE were added to the BEM solution. The mixture was cannulated into a 1 L 5-neck flask with standard set-up. The cylinder was rinsed with 20 ml heptane. The solution was stirred at room temperature at 150 rpm.

27.18 g 2-Et-HexOH was weighed into a 50 ml measuring cylinder, and diluted to 50 ml with heptane. The alcohol solution was transferred to the 60 ml addition funnel. The cylinder was rinsed with 10 ml heptane. The alcohol solution was added to the 1 L flask at room temperature in 15 mins. The colorless solution was stirred at room temperature for 45 mins.

100 mmol of $\text{CITi}(\text{O}^i\text{Pr})_3$ was weighed into a 100 ml measuring cylinder and transferred to the 60 ml addition funnel (50 ml at a time). The solution was added dropwise to the 1 L flask at room temperature over 39 minutes. The addition funnel was rinsed with 10 ml heptane. The reaction mixture was stirred at room temperature for one hour.

34.0 g $\text{Ti}(\text{O}^i\text{Bu})_4$ was weighed into a 200 ml flask. 38.0 g TiCl_4 was weighed into a 50 ml measuring cylinder and diluted to 50 ml with heptane. The TiCl_4 solution was transferred to the 200 ml flask. The cylinder was rinsed with 90 ml heptane. The orange-brown reaction mixture was stirred at room temperature for 45 mins.

50 The $\text{TiCl}_4/\text{Ti}(\text{O}^i\text{Bu})_4$ solution was transferred to the 60 ml addition funnel on the 1 L and added dropwise to the flask at room temperature over 1 hr and 40 mins. The reaction mixture was stirred at room temperature for another hour and then was allowed to settle. The supernatant was decanted. The solid was washed twice with 200 ml portions of heptane, then reslurried in ca. 180 ml heptane.

9.5 g TiCl_4 was weighed into a 50 ml measuring cylinder and diluted to 60 ml with heptane. The TiCl_4 solution was

added dropwise at room temperature over 30 mins. The addition funnel was rinsed with 10 ml heptane. The reaction mixture was stirred at room temperature for another hour. A slurry sample was taken. A small portion of the slurry was transferred to a 100 ml flask, decanted, washed and dried *in vacuo* (1.7 g).

7.18 g TEAL (24.8% in heptane) was weighed into a 100 ml measuring cylinder and diluted to 60 ml with heptane. The TEAL solution was added dropwise at room temperature over 26 mins. The reaction mixture was stirred at room temperature for another hour and then was allowed to settle. The supernatant was decanted. The solid was washed four times with 200 ml portions of heptane. The solid was dried *in vacuo* at room temperature for 2 hrs. (16.1 g).

TABLE II

Example	Yield(g)	Productivity (g/g • hr)	% fines (<125 µm)	Catalyst Particle Size (microns)
8.	375	28,000	1.0	11.0
12.	253	33,000	0.6	13.0
13.	304	37,000	0.8	12.4
14.	352	34,000	19.1	9.8
15.	378	48,000	3.2	13.5
16.	483	60,000	2.2	13.5
17.	155	64,000	9.8	10.8
18.	418	57,000	3.6	11.4
19.	406	42,000	2.6	14.7
20.	392	45,000	1.8	12.6
21.	492	51,000	4.1	12.8
22.	405	48,000	1.8	18.6

Concentration of the $\text{TiCl}_4/\text{Ti}(\text{O}i\text{Bu})_4$ and the amount of TiCl_4 used for the second titration did not have any effect on catalyst performance and fluff morphology. Temperatures of the first titration which were higher or lower than room temperature resulted in smaller catalyst particle size, slower catalyst settling and more fines in the polymer fluff. Catalyst synthesis is preferably carried out at room (ambient) temperature. Heat treatment after the first titration had little effect on catalyst morphology. Addition of DIAE results in slightly more fines but significantly improves catalyst activity. The electron donor is believed to increase the integrity of and reduce fragmentation of the catalyst particles. Use of an alkyl aluminum such as TEAL during catalyst synthesis helps reduce viscosity of the magnesium alkoxide. Addition of an electron donor such as an ether, e.g., diisooamyl ether (DIAE), may be used to further reduce the viscosity of the alkyl magnesium and the magnesium alkoxide reaction product. The preactivation of the catalyst increases catalyst particle size and narrows particle size distribution.

The amount of fines in the fluff that were produced by the improved catalyst is much less than that produced by the comparative catalyst. Considering fluff particles smaller than 125 microns, the fluff produced by the comparative catalyst contain 3-5% of such particles, while the improved catalyst produced fluffs containing 1-2%.

Claims

1. A catalyst for polymerizing ethylene comprising:

- a soluble magnesium compound of magnesium dialkoxide of the general formula $\text{Mg}(\text{OR})_2$ where R is a hydrocarbyl or substituted hydrocarbyl of 1 to 20 carbon atoms;
- a chlorinating agent capable of exchanging one chlorine for one alkoxide;
- a first chlorinating/titanating agent;
- a second stronger chlorinating/titanating agent; and
- an organoaluminum compound preactivating agent.

2. The catalyst of Claim 1 wherein the soluble magnesium compound is a reaction product of an alkyl magnesium compound of the general formula MgRR' , where R and R' are alkyl groups of 1-10 carbon atoms and may be the

same or different, and an alcohol of the general formula $R''OH$ where R'' is an alkyl group of 2-10 carbon atoms.

3. The catalyst of Claim 1 wherein the alkyl magnesium compound is magnesium diethyl, magnesium dipropyl, magnesium dibutyl or butylethylmagnesium.
- 5 4. The catalyst of Claim 1 wherein the alkyl magnesium compound is butylethylmagnesium.
5. The catalyst of Claim 1 wherein the alcohol is linear or branched.
- 10 6. The catalyst of Claim 1 wherein the alcohol is ethanol, propanol, isopropanol, butanol, isobutanol or 2-ethylhexanol.
7. The catalyst of Claim 6 wherein the alcohol is 2-ethylhexanol.
8. The catalyst of Claim 1 wherein the soluble magnesium compound is magnesium di(2-ethylhexoxide).
- 15 9. The catalyst of any one of Claims 1 to 8 additionally comprising an aluminum alkyl in a).
10. The catalyst of Claim 9 wherein the aluminum alkyl is triethylaluminum.
- 20 11. The catalyst of Claim 9 wherein the ratio of alkyl aluminum to magnesium is 0.001:1 to 1:1.
12. The catalyst of Claim 11 wherein the ratio of alkyl aluminum to magnesium is 0.01:1 to 0.1:1.
13. The catalyst of Claim 12 wherein the ratio of alkyl aluminum to magnesium is 0.03:1 to 0.05:1.
- 25 14. The catalyst of any one of Claims 1 to 13 additionally comprising an electron donor.
15. The catalyst of Claim 14 wherein the electron donor is diisoamyl ether.
- 30 16. The catalyst of Claim 14 wherein the ratio of the electron donor to magnesium is 0:1 to 10:1.
17. The catalyst of Claim 16 wherein the ratio of the electron donor to magnesium is 0.1:1 to 1:1.
18. The catalyst of any one of Claims 1 to 17 wherein the chlorinating agent is of the general formula $ClAR'''_x$ or $ClAOR'''_x$, where A is a nonreducing oxyphilic compound which is capable of exchanging one chloride for an alkoxide, R''' is alkyl and x is the valence of A minus 1.
- 35 19. The catalyst of Claim 18 wherein A is titanium, silicon, aluminum, carbon, tin or germanium.
- 40 20. The catalyst of Claim 19 wherein A is titanium or silicon and x is 3.
21. The catalyst of Claim 18 wherein R''' has 2-6 carbon atoms.
22. The catalyst of Claim 21 wherein R''' is methyl, ethyl, propyl, or isopropyl.
- 45 23. The catalyst of Claim 18 wherein the chlorinating agent is $ClTi(O^iPr)_3$ or $ClSi(Me)_3$.
24. The catalyst of Claim 23 wherein the chlorinating agent is $ClTi(O^iPr)_3$.
- 50 25. The catalyst of any one of Claims 1 to 24 wherein the first chlorinating/titanating agent is selected from the groups consisting of:
 - i) a blend of two tetra-substituted titanium compounds with all four substituents being the same and the substituents being a halide or an alkoxide or phenoxide with 2 to 10 carbon atoms;
 - 55 ii) a titanium alkoxyhalide of the general formula $Ti(OR)_aX_b$, where OR and X are an alkoxide and halide, respectively, $a + b$ is the valence of titanium; and
 - iii) a mixture of i) and ii)

26. The catalyst of Claim 25 wherein the first chlorinating/titanating agent is a blend of a titanium halide and a titanium alkoxide.
27. The catalyst of Claim 26 wherein the first chlorinating/titanating agent is a blend of TiCl_4 and $\text{Ti}(\text{OBu})_4$.
- 5 28. The catalyst of Claim 27 wherein the first chlorinating/titanating agent is a blend of TiCl_4 and $\text{Ti}(\text{OBu})_4$ in a range from 0.5:1 to 6:1 $\text{TiCl}_4/\text{Ti}(\text{OBu})_4$.
- 10 29. The catalyst of Claim 28 wherein the first chlorinating/titanating agent is a blend of TiCl_4 and $\text{Ti}(\text{OBu})_4$ in a range from 2:1 to 3:1.
30. The catalyst of Claim 29 wherein the ratio of titanium to magnesium is 3:1.
- 15 31. The catalyst of Claim 25 wherein the first chlorinating/titanating agent is $\text{Ti}(\text{OC}_2\text{H}_5)_3\text{Cl}$, $\text{Ti}(\text{OC}_3\text{H}_7)_2\text{Cl}_2$, $\text{Ti}(\text{OC}_3\text{H}_7)_3\text{Cl}$, $\text{Ti}(\text{OC}_4\text{H}_9)\text{Cl}_3$, $\text{Ti}(\text{OC}_6\text{H}_{13})_2\text{Cl}_2$, $\text{Ti}(\text{OC}_2\text{H}_5)_2\text{Br}_2$ or $\text{Ti}(\text{OC}_{12}\text{H}_{25})\text{Cl}_3$.
32. The catalyst of any one of Claims 1 to 31 additionally comprising an electron donor of the general formula $\text{RSi}(\text{OR}')_3$ where R and R' are alkyls with 1-5 carbon atoms and may be the same or different.
- 20 33. The catalyst of Claim 32 wherein the electron donor is methylsilyltriethoxide.
34. The catalyst of any one of Claims 1 to 33 wherein the second stronger chlorinating/titanating agent is a titanium halide.
- 25 35. The catalyst of Claim 34 wherein the second stronger chlorinating/titanating agent is titanium tetrachloride.
36. The catalyst of any one of Claims 1 to 35 wherein the range of titanium to magnesium is 0:1 to 2:1.
37. The catalyst of Claim 36 wherein the range of titanium to magnesium is 0.5:1 to 1:1.
- 30 38. The catalyst of any one of Claims 1 to 37 wherein the organoaluminum preactivating agent is an aluminum alkyl of the formula AlR^{A}_3 where R^{A} is an alkyl having 1-8 carbon atoms or a halide, R^{A} being the same or different and at least one R^{A} is an alkyl.
- 35 39. The catalyst of Claim 38 wherein the organoaluminum preactivating agent is a trialkyl aluminum.
40. The catalyst of Claim 39 wherein the organoaluminum preactivating agent is trimethyl aluminum, triethyl aluminum or triisobutyl aluminum.
- 40 41. The catalyst of Claim 40 wherein the organoaluminum preactivating agent is triethyl aluminum.
42. The catalyst of any one of Claims 1 to 41 wherein the ratio of Al to titanium is in the range from 0.1:1 to 2:1.
43. The catalyst of Claim 42 wherein the ratio of Al to titanium is in the range from 0.5:1 to 1.2:1.
- 45 44. The catalyst of Claim 43 wherein the ratio of Al to titanium is approximately 0.8:1.
45. A process for synthesizing a catalyst comprising:
 - 50 a) selecting a soluble magnesium dialkoxide compound of the general formula $\text{Mg}(\text{OR})_2$ where R is a hydrocarbyl or substituted hydrocarbyl of 1 to 20 carbon atoms;
 - b) adding a chlorinating agent capable of exchanging one chlorine for one alkoxide;
 - c) adding a first chlorinating/titanating agent;
 - d) adding a second stronger chlorinating/titanating agent;
 - 55 e) adding an organoaluminum preactivating agent.
46. The process of Claim 45 wherein the soluble magnesium compound of step a) is as described in any one of Claims 2 to 8.

47. The process of Claims 45 or 46 additionally selecting in step a) an aluminum alkyl as described in any one of Claims 10 to 13.
- 5 48. The process of any one of Claims 45 to 47 additionally comprising the step of adding an electron donor as described in any one of Claims 15 to 17.
49. The process of any one of Claims 45 to 48 wherein the chlorinating agent of step b) is as described in any one of Claims 18 to 24.
- 10 50. The process of any one of Claims 45 to 49 wherein the first chlorinating/titanating agent of step c) is as described in any one of Claims 25 to 31.
51. The process of any one of Claims 45 to 50 additionally comprising the step of adding an electron donor as described in Claims 33 or 34.
- 15 52. The process of any one of Claims 45 to 51 wherein the second stronger chlorinating/titanating agent of step d) is as described in any one of Claims 32 to 37.
53. The process of any one of Claims 45 to 52 wherein the organoaluminum preactivating agent of step e) is as described in Claims 38 to 44.
- 20 54. A process for polymerizing ethylene comprising:
 - a) providing a catalyst according to any one of Claims 1 to 44
 - 25 b) activating the catalyst with an organoaluminum compound
 - c) contacting the catalyst with ethylene monomer under polymerization conditions,
 - d) extracting polyethylene.
55. The process of Claim 54 additionally comprising prepolymerizing the catalyst.
- 30 56. The process of Claim 54 wherein polymerization is in bulk, slurry or gas phase.
57. The process of Claim 56 wherein polymerization is in slurry phase.
- 35 58. The process of Claim 57 wherein slurry polymerization is in a temperature range of 50-120°C.
59. The process of Claim 58 wherein slurry polymerization is in a temperature range of 50-100°C.
60. The process of Claim 59 wherein slurry polymerization is in a temperature range of 70-80°C.
- 40 61. The process of Claim 57 wherein slurry polymerization is in a pressure range of 3.45 to 55.16 bars (50-800 psi).
62. The process of Claim 61 wherein slurry polymerization is in a pressure range of 5.52 to 41.37 bars (80-600 psi).
- 45 63. The process of Claim 62 wherein slurry polymerization is in a pressure range of 6.89 to 10.34 bars (100-150 psi).
64. The process of Claim 54 wherein the olefin monomer is introduced into the polymerization reaction zone in a diluent which is a nonreactive heat transfer agent and is liquid at the reaction conditions.
- 50 65. The process of Claim 64 wherein the diluent is hexane or isobutane.
66. The process of Claim 57 wherein the slurry polymerization is a copolymerization of ethylene and butene.
67. The process of Claim 66 wherein butene is present at 0.01-20 mole percent.
- 55 68. The process of Claim 67 wherein butene is present at 0.02-1 mole percent.
69. The process of Claim 68 wherein butene is present from about 0.04 to about 0.08 mole percent.



European Patent
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EUROPEAN SEARCH REPORT

Application Number
EP 98 10 1439

DOCUMENTS CONSIDERED TO BE RELEVANT			
Category	Citation of document with indication, where appropriate, of relevant passages	Relevant to claim	CLASSIFICATION OF THE APPLICATION (Int.Cl.6)
X	EP 0 357 135 A (SHELL INT RESEARCH) * claims 4,6 * * example 1 *	1,25,34, 35,45, 50,52	C08F4/658 C08F10/00
A	EP 0 068 256 A (HOECHST AG) * example 1 *	1,25,34, 35,38, 45, 52-54, 56-59, 61,62	
A	WO 84 04925 A (DOW CHEMICAL CO) * example 4 *	1-3,5,6, 9,11,45, 54, 56-59, 61,64,65	
A	DE 34 32 759 A (BASF AG) * example 1 *	1,25-27, 34,35,45	TECHNICAL FIELDS SEARCHED (Int.Cl.6) C08F
The present search report has been drawn up for all claims			
Place of search THE HAGUE		Date of completion of the search 13 May 1998	Examiner Fischer, B
CATEGORY OF CITED DOCUMENTS X : particularly relevant if taken alone Y : particularly relevant if combined with another document of the same category A : technological background O : non-written disclosure P : intermediate document		T : theory or principle underlying the invention E : earlier patent document, but published on, or after the filing date D : document cited in the application L : document cited for other reasons & : member of the same patent family, corresponding document	

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